



Standard Test Methods for Arsenic in Water¹

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1. Scope

1.1 These test methods² cover the photometric and atomic absorption determination of arsenic in most waters and wastewaters. Three test methods are given as follows:

	Concentration Range	Sections
Test Method A—Silver Diethyldithiocarbamate Colorimetric	5 to 250 $\mu\text{g/L}$	7 to 15
Test Method B—Atomic Absorption, Hydride Generation	1 to 20 $\mu\text{g/L}$	16 to 24
Test Method C—Atomic Absorption, Graphite Furnace	5 to 100 $\mu\text{g/L}$	25 to 33

1.2 The analyst should direct attention to the precision and bias statements for each test method. It is the user's responsibility to ensure the validity of these test methods for waters of untested matrices.

1.3 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.* For specific hazard statements, see Note 1 and Note 5.

2. Referenced Documents

2.1 ASTM Standards:

- D 1129 Terminology Relating to Water³
- D 1193 Specification for Reagent Water³
- D 2777 Practice for Determination of Precision and Bias of Applicable Methods of Committee D-19 on Water³
- D 3370 Practices for Sampling Water from Closed Conduits³
- D 3919 Practice for Measuring Trace Elements in Water by Graphite Furnace Atomic Absorption Spectrophotometry³
- D 4841 Practice for Estimation of Holding Time for Water Samples Containing Organic and Inorganic Constituents³
- E 60 Practice for Photometric and Spectrophotometric

¹ These test methods are under the jurisdiction of ASTM Committee D-19 on Water and are the direct responsibility of Subcommittee D 19.05 on Inorganic Constituents in Water.

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² Similar to that appearing in *Standard Methods for the Examination of Water and Wastewater*, 12th edition, APHA, Inc., New York, NY, 1965, and identical with that in Brown, Eugene, Skougstad, M. W., and Fishman, M. J., "Methods for Collection and Analysis of Water Samples for Dissolved Minerals and Gases," *Techniques of Water-Resources Investigations of the U.S. Geological Survey*, Book 5, Chapter , 1970 p. 46.

³ *Annual Book of ASTM Standards*, Vol 11.01.

Methods for Chemical Analysis of Metals⁴

E 275 Practice for Describing and Measuring Performance of Ultraviolet, Visible, and Near Infrared Spectrophotometers⁵

3. Terminology

3.1 Definitions of Term Specific to This Standard:

3.1.1 For definitions of terms used in these test methods refer to Terminology D 1129.

3.2 Definitions of Terms Specific to This Standard:

3.2.1 *total recoverable arsenic*—an arbitrary analytical term relating to the forms of arsenic that are determinable by the digestion method which is included in the procedure. Some organic-arsenic compounds, such as phenylarsonic acid, disodium methane arsonate, and dimethylarsonic acid, are not recovered completely during the digestion step.

4. Significance and Use

4.1 Herbicides, insecticides, and many industrial effluents contain arsenic and are potential sources of water pollution. Arsenic is significant because of its adverse physiological effects on humans.

5. Purity of Reagents

5.1 Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society, where such specifications are available.⁶ Other grades may be used, provided it is first ascertained that the reagent is of sufficiently high purity to permit its use without lessening the accuracy of the determination.

5.2 *Purity of Water*— Unless otherwise indicated, references to water shall be understood to mean reagent water conforming to Specification D 1193 Type I. Other reagent water types may be used provided it is first ascertained that the water is of sufficiently high purity to permit its use without adversely affecting the bias and precision of the test method.

⁴ *Annual Book of ASTM Standards*, Vol 03.05.

⁵ *Annual Book of ASTM Standards*, Vol 03.06.

⁶ *Reagent Chemicals, American Chemical Society Specifications*, American Chemical Society, Washington, DC. For suggestions on the testing of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K., and the *United States Pharmacopeia and National Formulary*, U.S. Pharmaceutical Convention, Inc. (USPC), Rockville, MD.

Type II water was specified at the time of round robin testing of this test method.

6. Sampling

6.1 Collect the sample in accordance with Practices D 3370.

6.2 Preserve the samples with HNO_3 (sp gr 1.42) to a pH of 2 or less immediately at the time of collection; normally about 2 mL/L is required. If only dissolved arsenic is to be determined, filter the sample through a 0.45- μm membrane filter before acidification. The holding times for the samples may be calculated in accordance with Practice D 4841.

TEST METHOD A—SILVER DIETHYLDITHIOCARBAMATE COLORIMETRIC

7. Scope

7.1 This test method covers the determination of dissolved and total recoverable arsenic in most waters and waste waters in the range from 5 to 250 $\mu\text{g/L}$ of arsenic.

7.2 The precision and bias data were obtained on reagent water, river water, and process water. The information on precision and bias may not apply to other waters. It is the user's responsibility to ensure the validity of this test method for waters of untested matrices.

8. Summary of Test Method

8.1 Organic arsenic-containing compounds are decomposed by adding sulfuric and nitric acids and repeatedly evaporating the sample to fumes of sulfur trioxide. The arsenic (V) so produced, together with inorganic arsenic originally present, is subsequently reduced to arsenic (III) by potassium iodide and stannous chloride, and finally to gaseous arsine by zinc in hydrochloric acid solution. The resulting mixture of gases is passed through a scrubber containing borosilicate wool impregnated with lead acetate solution and then into an absorption tube containing a solution of silver diethyldithiocarbamate in pyridine. Arsine reacts with this reagent to form a red-colored silver sol having maximum absorbance at about 540 nm. The absorbance of the solution is measured photometrically and the arsenic determined by reference to an analytical curve prepared from standards.

9. Interferences

9.1 Although many samples are relatively free of interferences, several metals, notably cobalt, nickel, mercury, silver, platinum, copper, chromium, and molybdenum, may interfere with the evolution of arsine and with the recovery of arsenic. The presence of any or all of these metals in a sample being analyzed must be considered as a potential source of interference, and the analyst must fully determine the extent of actual interference, if any. This could be accomplished by spiking.

9.2 Hydrogen sulfide and other sulfides interfere, but commonly encountered quantities are effectively removed by the lead acetate scrubber and the digestion.

9.3 Antimony interferes by forming stibine, which distills along with the arsine. Stibine reacts with the color-forming reagent to form a somewhat similar red sol having maximum absorbance near 510 nm. The sensitivity for antimony at 540 nm is only about 8 % that of arsenic (1 mg/L of antimony will show an apparent presence of 0.08 mg/L of arsenic).

9.4 Nitric acid interferes with the test and must be completely eliminated during the digestion.

10. Apparatus

10.1 *Arsine Generator, Scrubber, and Absorber*⁷, assembled as shown in Fig. 1.

10.2 *Spectrophotometer or Filter Photometer*, suitable for use at 540 nm and providing a light path of at least 10 mm. The filter photometer and photometric practice prescribed in this method shall conform to Practice E 60. The spectrophotometer shall conform to Practice E 275.

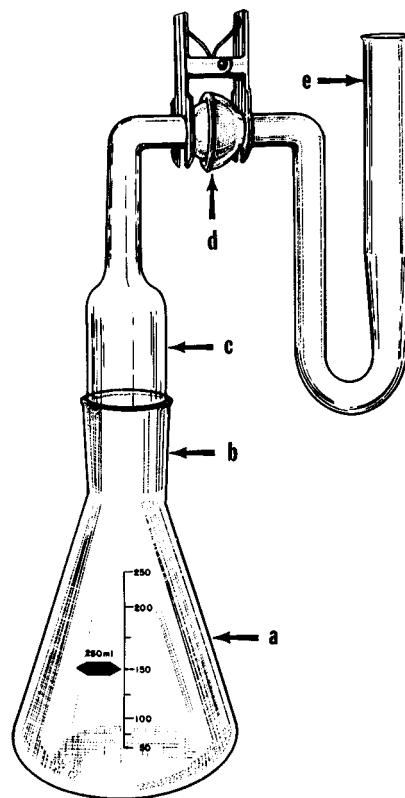
11. Reagents

11.1 *Arsenic Solution, Stock* (1.00 mL = 1.00 mg As)—Dissolve 1.320 g of arsenic trioxide (As_2O_3) (**Warning**, see Note 1), dried for at least 1 h at 110°C, in 10 mL of NaOH solution (420 g/L) and dilute to 1 L with water. This solution is stable.

NOTE 1—**Warning:** Arsenic trioxide is extremely toxic. Avoid ingestion or inhalation of dry powder during standard preparation. Wash hands thoroughly immediately after handling arsenic trioxide. Under no circumstances pipet any arsenic solutions by mouth.

11.2 *Arsenic Solution, Intermediate* (1.00 mL = 10.0 μg As)—Dilute 5.00 mL of arsenic stock solution to 500 mL with water.

⁷ Available commercially.



- (a) Generator flask, borosilicate glass, 250-mL capacity.
- (b) Standard-taper neck 24/40.
- (c) Scrubber, borosilicate glass wool impregnated with lead acetate.
- (d) Ground-glass ball-and-socket joint.
- (e) Absorber: add AgDDC solution and pack with glass beads.

FIG. 1 Arsine Generator, Scrubber, and Absorber⁷

11.3 *Arsenic Solution, Standard* (1.00 mL = 1.00 µg As)—Dilute 10.0 mL of arsenic intermediate solution to 100 mL with water. Prepare fresh before each use.

11.4 *Hydrochloric Acid* (sp gr 1.19)—Concentrated hydrochloric acid (HCl). Use analytical grade acid with an arsenic content not greater than 1×10^{-6} %.

11.5 *Lead Acetate Solution* (100 g/L)—Dissolve 10 g of lead acetate ($\text{Pb}(\text{C}_2\text{H}_3\text{O}_2)_2 \cdot 3\text{H}_2\text{O}$) in 100 mL of water. Store reagent in a tightly stoppered container.

11.6 *Nitric Acid* (sp gr 1.42)—Concentrated nitric acid (HNO_3). Use analytical grade acid with an arsenic content not greater than 1×10^{-6} %.

11.7 *Nitric Acid* (1 + 1)—Add 250 mL of concentrated nitric acid (sp gr 1.42) to 250 mL of water.

11.8 *Potassium Iodide Solution* (150 g/L)—Dissolve 15 g of potassium iodide (KI) in 100 mL of water. Store in an amber bottle.

11.9 *Silver Diethyldithiocarbamate Solution*—Dissolve 1 g of silver diethyldithiocarbamate (AgDDC) in 200 mL of pyridine. This solution is stable for at least several months when stored in an amber bottle.

11.10 *Sodium Hydroxide Solution* (420 g/L)—Dissolve 42 g of sodium hydroxide (NaOH) pellets in 100 mL of water.

11.11 *Stannous Chloride Solution*—Dissolve 40 g of arsenic-free stannous chloride ($\text{SnCl}_2 \cdot 2\text{H}_2\text{O}$) in 100 mL of HCl (sp gr 1.19). Add a few small pieces of mossy tin.

11.12 *Sulfuric Acid* (1 + 1)—Cautiously, and with constant stirring and cooling, add 250 mL of concentrated H_2SO_4 (sp gr 1.84) to 250 mL of water.

11.13 *Zinc, Granular*, 20-mesh. Arsenic content must not exceed 1×10^{-6} %.

12. Standardization

12.1 Clean all glassware before use by rinsing first with hot HNO_3 (1 + 1) and then with water. The absorbers must be additionally rinsed with acetone and then air-dried.

12.2 Prepare, in a 250-mL generator flask, a blank and sufficient standards containing from 0.0 to 25.0 µg of arsenic by diluting 0.0 to 25.0-mL portions of the arsenic standard solution to approximately 100 mL with water.

12.3 Proceed as directed in 13.3-13.9.

12.4 Construct an analytical curve by plotting the absorbances of standards versus micrograms of arsenic.

NOTE 2—The response is linear up to 15 µg of arsenic; however, because the curve is nonlinear above 15 µg, it is necessary to have sufficient standards above 15 µg to permit constructing an accurate curve.

13. Procedure

13.1 Clean all glassware before use by rinsing first with hot HNO_3 (1 + 1) and then with water. The absorbers must be additionally rinsed with acetone and then air-dried.

13.2 Pipet a volume of well-mixed acidified sample containing less than 25 µg of arsenic (100 mL maximum) into a generating flask and dilute to approximately 100 mL.

NOTE 3—If only dissolved arsenic is to be determined use a filtered and acidified sample (see 6.2).

13.3 To each flask, add 7 mL of H_2SO_4 (1 + 1) and 5 mL of concentrated HNO_3 . Add a small boiling chip and carefully

evaporate to dense fumes of SO_3 , maintaining an excess of HNO_3 until all organic matter is destroyed. This prevents darkening of the solution and possible reduction and loss of arsenic. Cool, add 25 mL of water, and again evaporate to dense fumes of SO_3 . Maintain heating for 15 min to expel oxides of nitrogen.

13.4 Cool, and adjust the volume in each flask to approximately 100 mL with water.

13.5 To each flask add successively, with thorough mixing after each addition, 8 mL of concentrated HCl, 4 mL of KI solution, and 1 mL of SnCl_2 solution. Allow about 15 min for complete reduction of the arsenic to the trivalent state.

13.6 Place in each scrubber a plug of borosilicate wool that has been impregnated with lead acetate solution. Assemble the generator, scrubber, and absorber, making certain that all parts fit and are correctly adjusted. Add 3.00 mL of silver diethyldithiocarbamate-pyridine solution to each absorber. Add glass beads to the absorbers until the liquid just covers them.

NOTE 4—Four millilitres of silver diethyldithiocarbamate-pyridine solution may be used with some loss of sensitivity.

13.7 Disconnect each generator, add 6 g of zinc, and reconnect immediately.

13.8 Allow 30 min for complete evolution of arsine. Warm the generator flasks for a few minutes to make sure that all arsine is released.

13.9 Pour the solutions from the absorbers directly into clean spectrophotometer cells and within 30 min measure the absorbance of each at 540 nm.

14. Calculation

14.1 Determine the weight of arsenic in each sample by referring to the analytical curve. Calculate the concentration of arsenic in the sample in micrograms per litre, using Eq 1:

$$\text{Arsenic, } \mu\text{g/L} = 1000 \text{ W/V} \quad (1)$$

where:

V = volume of sample, mL, and

W = weight of arsenic in sample, µg.

15. Precision and Bias ⁸

15.1 The single-operator and overall precision of this method for three laboratories, which included a total of six operators analyzing each sample on three different days, within its designated range varies with the quantity being tested in accordance with Table 1.

15.2 Recoveries of known amounts of arsenic (arsenic trioxide) in a series of prepared standards are given in Table 1.

15.3 The precision and bias data were obtained on reagent water, river water, and process water. The information on precision and bias may not apply to other waters. It is the user's responsibility to ensure the validity of this test method for waters of untested matrices.

15.4 Three independent laboratories participated in the roundrobin study. Precision and bias for this test method

⁸ Supporting data are available from ASTM Headquarters. Request RR:D 19 - 1049.

TABLE 1 Precision and Bias for Arsenic by Test Method A, Diethyldithiocarbamate Colorimetric

Water	Amount Added, µg/L	Amount Found, µg/L	S_r	S_o	Bias, %
Reagent Type II	25.0	23.66	1.76	1.78	-5.4
	100.0	95.28	5.21	5.24	-4.7
	200.0	194.99	8.43	8.79	-2.6
Water of Choice	25.0	24.76	2.07	1.84	-0.96
	100.0	97.00	4.15	3.78	-3.0
	200.0	189.01	9.96	9.70	-5.5

conform to Practice D 2777 – 77, which was in place at the time of collaborative testing. Under the allowances made in 1.5 of Practice D 2777 – 86, these precision and bias data do meet existing requirements for interlaboratory studies of Committee D-19 test methods.

TEST METHOD B—ATOMIC ABSORPTION, HYDRIDE GENERATION

16. Scope

16.1 This test method covers the determination of dissolved and total recoverable arsenic in most waters and wastewaters in the range from 1 to 20 µg/L of arsenic. The range may be extended by dilution of the sample.

16.2 The precision and bias data were obtained on reagent water, tap water, salt water, river water, and untreated wastewater. The information on precision and bias may not apply to other waters. It is the user’s responsibility to ensure the validity of this test method for waters of untested matrices.

17. Summary of Test Method

17.1 Organic arsenic-containing compounds are decomposed by adding sulfuric and nitric acids and repeatedly

evaporating the sample to fumes of sulfur trioxide. The arsenic (V) so produced, together with inorganic arsenic originally present, is subsequently reduced to arsenic (III) by potassium iodide and stannous chloride, and finally to gaseous arsine by zinc in hydrochloric acid solution. Alternatively, the arsenic is converted to arsine by sodium borohydride in hydrochloric acid solution. The arsine is removed from solution by aeration and swept by a flow of nitrogen into a hydrogen flame where it is determined by atomic absorption at 193.7 nm.

18. Interferences

18.1 See 9.1.

19. Apparatus

19.1 *Arsine Vapor Analyzer*, assembled as shown in Fig. 2⁹.

19.2 *Atomic Absorption Spectrophotometer*, (**Warning**, see Note 5) for use at 193.7 nm.

NOTE 5—**Warning:** Because of the toxicity of arsenic, a well-ventilated hood must be used with the atomic absorption spectrometer.

NOTE 6—Follow the manufacturer’s instructions for all instrumental parameters.

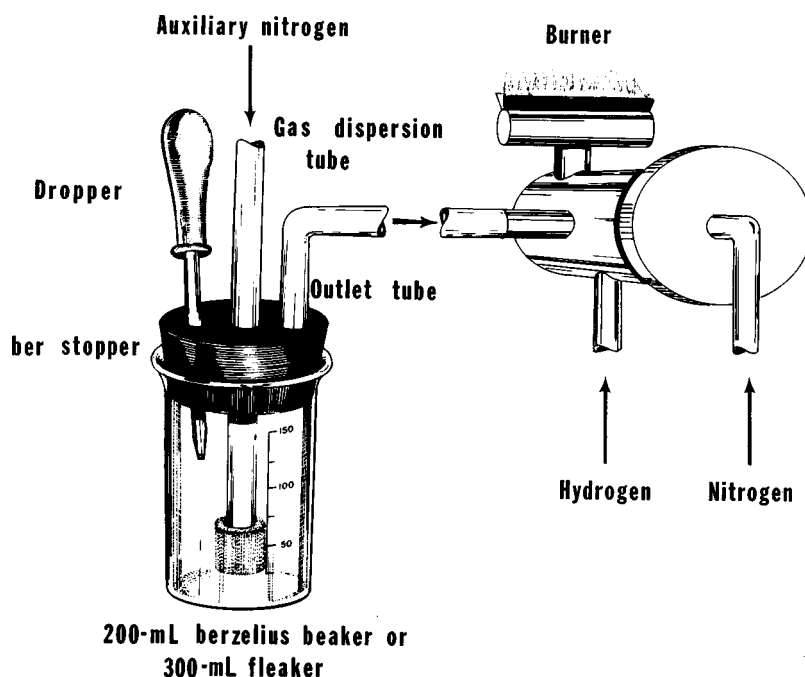
19.2.1 *Arsenic Light Source*—Arsenic electrodeless discharge lamp or hollow-cathode lamp.

20. Reagents and Materials

20.1 *Arsenic Solution, Stock* (1.00 mL = 1.00 mg As)—See 11.1.

20.2 *Arsenic Solution, Intermediate* (1.00 mL = 10.0 µg As)—See 11.2.

⁹ A static system, such as one using a balloon, has been found satisfactory for this purpose. See McFarren, E. F., “New Simplified Methods for Metal Analysis,” *Journal of American Water Works Association*, Vol 64, 1972, p. 28.



NOTE 1—Fleaker, trademarked product of Corning Glass Works, and Berzelius beaker are available from most laboratory apparatus dealers.

FIG. 2 Arsine Vapor Analyzer

20.3 *Arsenic Solution, Standard* (1.00 mL = 0.10 µg As)—Dilute 10.0 mL of arsenic intermediate solution to 1000 mL with water. Prepare fresh before each use.

20.4 *Hydrochloric Acid* (sp gr 1.19)—See 11.4.

20.5 *Nitric Acid* (sp gr 1.42)—See 11.6.

20.6 *Nitric Acid* (1 + 1)—See 11.7.

20.7 *Nitric Acid* (1 + 4)—Add 20 mL of nitric acid (sp gr 1.42) to 80 mL of water.

20.8 *Potassium Iodide Solution* (150 g/L)—See 11.8.

20.9 *Sodium Borohydride Solution* (4 g/100 mL)—Dissolve 4 g of sodium borohydride (NaBH₄) in 100 mL of water. Prepare fresh before each use.

20.10 *Stannous Chloride Solution* (400 g/L)—See 11.11.

20.11 *Sulfuric Acid* (1 + 1)—See 11.12.

20.12 *Zinc Metal (Dust) Suspension*—Add 10 g of zinc dust to 20 mL of water.

20.13 *Hydrogen*—Set burner control box to a gage pressure of 8 psi (55 kPa) and adjust the flowmeter to approximately 6 L/min.

20.14 *Nitrogen or Argon*—Set burner control box to a gage pressure of 30 psi (207 kPa) and adjust the flowmeter for maximum sensitivity by volatilizing standards. A flow of approximately 8 L/min has been found satisfactory for this purpose. This will depend on the burner used.

21. Standardization

21.1 Clean all glassware before use by rinsing first with hot HNO₃(1 + 1) and then with water.

21.2 Prepare, in 200-mL Berzelius beakers or similar apparatus, a blank and sufficient standards containing from 0.0 to 1.0 µg of arsenic by diluting 0.0 to 10.0-mL portions of the arsenic standard solution to approximately 50 mL.

21.3 Proceed as directed in 22.1.3-22.1.8 or 22.2.3-22.2.7.

21.4 Prepare an analytical curve by plotting recorder scale readings versus micrograms of arsenic on linear graph paper. Alternatively, read directly in concentration if a concentration readout is provided with the instrument.

22. Procedure

22.1 Determination of Arsenic with Zinc:

22.1.1 Clean all glassware before use by rinsing first with hot HNO₃(1 + 1) and then with water.

22.1.2 Pipet a volume of well-mixed acidified sample containing less than 1.0 µg of arsenic (50-mL maximum) into a 200-mL Berzelius beaker (or similar apparatus) and dilute to approximately 50 mL.

NOTE 7—If only dissolved arsenic is to be determined use a filtered and acidified sample (see 6.2).

22.1.3 To each beaker, add 7 mL of H₂SO₄(1 + 1) and 5 mL of concentrated HNO₃(sp gr 1.42). Add a small boiling chip and carefully evaporate to fumes of SO₃, maintaining an excess of HNO₃ until all organic matter is destroyed. This prevents darkening of the solution and possible reduction and loss of arsenic. Cool, add 25 mL of water, and again evaporate to fumes of SO₃ to expel oxides of nitrogen.

22.1.4 Cool, and adjust the volume in each beaker to approximately 50 mL with water.

22.1.5 To each beaker, add successively, with thorough mixing after each addition, 8 mL of HCl (sp gr 1.19), 5 mL of

KI solution, and 1 mL of SnCl₂ solution. Allow about 15 min for reduction of the arsenic to the trivalent state.

22.1.6 Attach one beaker at a time to the rubber stopper containing the gas dispersion tube.

22.1.7 Fill the dropper or syringe with 2 mL of zinc dust suspension and insert into the hole in the rubber stopper.

NOTE 8—The zinc dust is kept in suspension by continuous stirring. A magnetic stirrer is satisfactory.

22.1.8 Add the zinc suspension to the sample solution. After the absorbance has reached a maximum and has returned to the baseline remove the beaker. Rinse the gas dispersion tube first in HNO₃ (1 + 4), and then in water before proceeding to the next sample. Treat each succeeding sample, blank, and standard in a like manner.

22.2 Determination of Arsenic with Sodium Borohydride:

22.2.1 Clean all glassware before use by rinsing first with hot HNO₃(1 + 1) and then with water.

22.2.2 Pipet a volume of well-mixed acidified sample containing less than 1.0 µg of arsenic (50 mL maximum) into a 200-mL Berzelius beaker (or similar apparatus) and dilute to approximately 50 mL (see Note 7).

22.2.3 To each beaker, add 7 mL of H₂SO₄(1 + 1) and 5 mL of concentrated HNO₃. Add a small boiling chip and carefully evaporate to fumes of SO₃, maintaining an excess of HNO₃ until all organic matter is destroyed. This prevents darkening of the solution and possible reduction and loss of arsenic. Cool, add 25 mL of water, and again evaporate to fumes of SO₃ to expel oxides of nitrogen.

22.2.4 Cool, and adjust the volume in each beaker to approximately 50 mL with water.

22.2.5 Add 8 mL of concentrated HCl and mix.

22.2.6 Attach one beaker at a time to the rubber stopper containing the gas dispersion tube.

22.2.7 Fill the dropper or syringe with 0.5 mL of sodium borohydride solution and insert into the hole in the rubber stopper.

22.2.8 Add the sodium borohydride solution to the sample solution. After the absorbance has reached a maximum and has returned to the baseline remove the beaker. Rinse the gas dispersion tube with water before proceeding to the next sample. Treat each succeeding sample, blank, and standard in a like manner.

23. Calculation

23.1 Determine the weight or concentration of arsenic in each sample by referring to 21.4. If the weight is determined from the analytical curve, calculate the concentration of arsenic in the sample in micrograms per litre, using Eq 2:

$$\text{Arsenic } \mu\text{g/L} = 1000 W/V \quad (2)$$

where:

V = volume of sample, mL, and

W = weight of arsenic in sample, µg.

24. Precision and Bias¹⁰

24.1 The single-operator and overall precision of this test

¹⁰ Supporting data are available from ASTM Headquarters. Request RR:D19-1050.

method for six laboratories, which included a total of ten operators analyzing each sample on three different days, within its designated range varies with the quantity being tested in accordance with Table 2.

24.2 See Table 2 for recoveries of known amounts of arsenic (arsenic trioxide) in a series of prepared standards.

24.3 The precision and bias data were obtained on reagent water, tap water, salt water, river water, and untreated wastewater. It is the user's responsibility to ensure the validity of this test method for waters of untested matrices.

TEST METHOD C—ATOMIC ABSORPTION, GRAPHITE FURNACE

25. Scope

25.1 This test method covers the determination of dissolved and total recoverable arsenic in most waters and wastewaters.

25.2 This test method is applicable in the range from 5 to 100 µg/L of arsenic using a 20-µL injection. The range can be increased or decreased by varying the volume of sample injected or the instrumental settings. High concentrations may be diluted but preferably should be analyzed by the atomic absorption-hydride method.

25.3 This test method has been used successfully with reagent water, lake water, river water, well water, filtered well water, and condensate from a medium Btu coal gasification process. It is the user's responsibility to ensure the validity of the test method to other matrices.

25.4 The analyst is encouraged to consult Practice D 3919 for a general discussion of interferences and sample analysis procedures for graphite furnace atomic absorption spectrophotometry.

26. Summary of Test Method

26.1 Arsenic is determined by an atomic-absorption spectrophotometer used in conjunction with a graphite furnace. A sample is placed in a graphite tube, evaporated to dryness, charred (pyrolyzed or ashed), and atomized. Since the graphite furnace uses the sample much more efficiently than flame atomization, the detection of low concentrations of elements in small sample volumes is possible. Finally, the absorption signal generated during atomization is recorded and compared to standards. A general guide for the application of the graphite furnace is given in Practice D 3919.

26.2 Dissolved arsenic is determined on a filtered and acidified sample with no pretreatment.

26.3 Total recoverable arsenic is determined following acid digestion and centrifugation. Because chlorides interfere with

furnace procedures for some metals, the use of hydrochloric acid in the digestion or solubilization step is to be avoided.

27. Interferences

27.1 For a complete discussion on general interferences with furnace procedures, the analyst is referred to Practice D 3919.

28. Apparatus

28.1 *Atomic-Absorption Spectrophotometer*, for use at 193.7 nm with background correction.

NOTE 9—A wavelength other than 193.7 nm may be used if it has been determined to be suitable. Greater linearity may be obtained at high concentrations by using a less sensitive wavelength.

NOTE 10—The manufacturer's instructions should be followed for all instrumental parameters.

28.2 *Centrifuge*, capable of holding centrifuge tubes of 15-mL volume.

28.3 *Centrifuge Tubes*, graduated centrifuge tubes of 15-mL capacity with stoppers.

28.4 *Graphite Furnace*, capable of reaching temperatures sufficient to atomize arsenic.

28.5 *Graphite Tubes*, compatible with the furnace device. Standard graphite tubes are recommended for the determination of arsenic.

28.6 *Light Source*—Arsenic electrodeless discharge lamps are recommended, but hollow-cathode lamps may be used.

28.7 *Pipets*—Microlitre with disposable tips. Sizes may range from 1 to 100 µL, as required.

28.8 *Strip-Chart Recorder*—A recorder is strongly recommended so that there will be a permanent record and any problems with the analysis can be easily recognized (such as drift, incomplete atomization, changes in sensitivity, etc.). A fast-response recorder (0.2 s or less for full-scale deflection) is recommended to ensure accuracy. Electronic peak-measuring devices have also been found useful.

28.9 *Automatic Sampling*, is recommended if available.

29. Reagents and Materials

29.1 *Arsenic Solution, Intermediate* (1.00 mL = 10.0 µg As)—See 11.2.

29.2 *Arsenic Solution, Standard* (1.00 mL = 1.00 µg As)—Dilute 10.0 mL of arsenic intermediate solution and 1 mL of HNO₃ (sp gr 1.42) to 100 mL. This standard is used to prepare working standards at the time of analysis.

29.3 *Hydrogen Peroxide* (30 %)—Hydrogen peroxide (H₂O₂).

29.4 *Nickel Nitrate Solution* (1.0 mL = 10 mg Ni)—Dissolve 5.0 g of nickel nitrate [Ni(NO₃)₂·6H₂O] in water and dilute to 100 mL.

29.5 *Nitric Acid* (sp gr 1.42)—Concentrated nitric acid (HNO₃).

29.6 *Nitric Acid* (1 + 9)—Add 1 mL of HNO₃ (sp gr 1.42) to 9 mL of water.

29.7 *Support Gas*—Prepurified argon is the usual support gas. Nitrogen may be used if recommended by the instrument manufacturer.

30. Standardization

30.1 Set the instrumental parameters to the manufacturer's

TABLE 2 Precision and Bias for Arsenic by Test Method B, Atomic Absorption-Hydride Generation

Water	Amount Added, µg/L	Amount Found, µg/L	S _r	S _o	Bias, %
Reagent Type II	3	3.16	0.76	0.74	+5
	10	9.74	0.93	0.97	-3
	18	17.67	1.81	1.93	-2
Water of Choice	3	2.70	0.70	0.48	-10
	10	8.76	1.93	0.94	-12
	18	18.07	2.93	2.22	+0.4

specifications. Follow the general instructions as provided in Practice D 3919.

30.2 Prepare a calibration curve using a blank and a series of standards in accordance with Practice D 3919.

NOTE 11—It is essential that the concentrations of the nickel nitrate and the nitric acid be equal for both standards and samples.

31. Procedure

31.1 Clean all glassware to be used for preparation of standard solutions or in the digestion step, or both, by rinsing first with HNO₃(1 + 1) and then with water. If possible, soak the glassware overnight in HNO₃(1 + 1).

31.2 If only dissolved arsenic is to be determined, add 8.0 mL of a filtered and acidified sample to a beaker or flask. Then add 1.0 mL of HNO₃(1 + 9) and 1.0 mL of nickel nitrate solution. Mix the solution thoroughly and proceed to 31.6.

31.3 For total arsenic, measure 30 mL of each standard and well-mixed sample to a 150-mL beaker. Add 0.25 mL of HNO₃(sp gr 1.42) and 2 mL of hydrogen peroxide (30 %) to the sample and mix thoroughly. Heat the samples at 95°C on a hot-plate or steam bath, in a well-ventilated fume hood, until the volume has been reduced to approximately 10 mL.

31.4 Cool and quantitatively transfer the sample to a 15-mL centrifuge tube. Dilute to volume with water, cap the tube, and mix the solution thoroughly. If undissolved material is present, centrifuge the sample for a few minutes to obtain a clear solution.

31.5 Pipet 5.0 mL of the supernatant liquid into a 10-mL volumetric flask. Add 1 mL of nickel solution and dilute to volume.

31.6 Inject a measured aliquot of sample into the furnace device following the directions as provided by the particular instrument manufacturer. Refer to Practice D 3919.

32. Calculation

32.1 Determine the concentration of arsenic in each sample by referring to Practice D 3919.

33. Precision and Bias ¹¹

33.1 The precision of this test method was tested by 12 laboratories in reagent water, lake water, river water, well water, filtered tap water, and condensate from a medium Btu coal gasification process. Two laboratories reported data from two operators. Although multiple injections may have been made, the report sheets provided allowed only for reporting single values. Thus, no single operator precision data can be calculated.

33.1.1 The overall precision of this test method, within its designated range for reagent water and selected water matrices, varies with the quantity tested as shown in Table 3.

33.1.2 Recovery and precision data for this test method are listed in Table 3.

33.2 The information on precision and bias may not apply to other waters. It is the user's responsibility to ensure the validity of this test method for waters of untested matrices.

34. Keywords

34.1 arsenic; atomic absorption; colorimetric (Test Method A); graphite furnace (Test Method C); hydride (Test Method B); water

¹¹ Supporting data are available from ASTM Headquarters. Request RR:D19-1108.

TABLE 3 Precision and Bias for Arsenic by Test Method C, Atomic Absorption-Graphite Furnace

Water	Amount Added, µg/L	Amount Found, µg/L	S _t	Bias, %
Reagent Type II	6.0	5.35	1.14	-11.0
	22.0	23.10	2.96	+ 5.0
	72.0	71.30	6.68	-1.0
Water of Choice	6.0	5.21	0.89	-13.0
	22.0	23.20	3.28	+ 5.4
	72.0	71.30	6.21	-1.0

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