



Designation: **E 1910 – 9700**

Standard Test Methods for Agricultural Acidifiers and Basic Blends¹

This standard is issued under the fixed designation E 1910; the number immediately following the designation indicates the year of original adoption or, in the case of revision, the year of last revision. A number in parentheses indicates the year of last reapproval. A superscript epsilon (ϵ) indicates an editorial change since the last revision or reapproval.

1. Scope

1.1 These test methods cover determining if an adjuvant meets the definitions of an acidifier or a basic blend in Terminology E 1519.

1.2 The values stated in both inch-pound and SI units are to be regarded separately as the standard. The values given in parentheses are provided for information only.

1.3 *This standard does not purport to address all of the safety concerns, if any, associated with its use. It is the responsibility of the user of this standard to establish appropriate safety and health practices and determine the applicability of regulatory limitations prior to use.*

2. Referenced Documents

2.1 *ASTM Standards:*

¹ These test methods are under the jurisdiction of ASTM Committee E35 on Pesticides and are the direct responsibility of Subcommittee E35.22 on Pesticide Formulation and Application Systems.

Current edition approved September 10, 1997; 2000. Published January 2001. Originally published as E 1910 – 987. Last previous edition E 1910 – 97.

~~D 1193 Specification 1126 Test Method for Reagent Hardness in Water~~²

~~D 1193 Specification for Reagent Water~~²

E 70 Test Method for pH of Aqueous Solutions With the Glass Electrode³

E 200 Practice for the Preparation, Standardization, and Storage of Standard and Reagent Solutions for Chemical Analysis³

E 1519 Terminology Relating to Agricultural Tank Mix Adjuvants⁴

3. Summary of Test Method

3.1 Test Method A measures the change in pH of a sample on the addition of an acidifier or a basic blend. Test Method B measures the ~~total acidity~~ buffer capacity of the adjuvant by titration.

4. Significance and Use

4.1 This test method is intended for products sold as tank mix adjuvants for use with agricultural products.

4.2 Part A provides a determination of whether the pH modification characteristics of the adjuvant tested meets the definition of an acidifier or a basic blend in Terminology E 1519. Part B provides a measure of the ~~acidification~~ buffer capacity of the adjuvant.

4.3 Other pH standards can be used (see Test Method E 70) as long as the pH's measured are not outside the range of the standards used.

4.4 If samples of the material to be used in a particular application are available, the first method can be used to determine what adjuvant rate is required to obtain the desired pH. The ability of an adjuvant to ~~acidify~~ adjust the pH of a particular system will depend on the initial pH before addition.

4.5 The second method can be used ~~for determining~~ to determine the amount of an adjuvant required to acidify ~~change the pH of a quantity of spray solution to the required value~~ (see Note 1).

NOTE 1—Many acidifiers ~~are and basic blends also are~~ buffers, so the pH change does not provide a useful measure of ~~acidification~~ buffer capacity.

5. Apparatus

5.1 *pH Meter*, accurate to 0.01 unit.

5.2 *Combination Glass/Calomel Electrode*, or separate glass and calomel electrodes or other suitable pH electrode system.

5.3 *Beakers*, 250 mL.

5.4 *Magnetic Stirrer*, 120 to 1200 rpm, or equivalent.

5.5 *Magnetic Stir Bar*, 1 in. (2.5 cm) in length and 3/8 in. (0.95 cm) in diameter.

5.6 *Burette*, 50 mL, Class A.

6. Reagents

6.1 *Purity of Reagents*—Reagent grade chemicals shall be used in all tests. Unless otherwise indicated, it is intended that all reagents shall conform to the specifications of the Committee on Analytical Reagents of the American Chemical Society where such specifications are available.

6.2 *Purity of Water*— Unless otherwise indicated, reference to water shall be understood to mean reagent water, Type IV, as defined by Specification D 1193 (see Note 2).

NOTE 2—Type IV grade reagent water may be prepared by distillation, ion exchange, reverse osmosis, electrodialysis, or a combination thereof.

6.3 *Synthetic Hard Water Stock*, transfer 12.14 g of anhydrous calcium chloride (CaCl₂) and 5.55 g of magnesium chloride hexahydrate (MgCl₂·6H₂O) to a 1000-mL volumetric flask. Dissolve the reagents with approximately 750 mL of water and equilibrate to 20°C. Dilute the solution to 1000 mL total volume with water at 20°C, stopper the flask and mix the solution thoroughly. This mixture has a hardness equivalent to 13680 ppm of calcium carbonate (CaCO₃) and is based on a compositional ratio of 4:1 calcium to magnesium.

6.4 *Standard Hard Water*, equivalent to a total hardness of 342 ppm of calcium carbonate (CaCO₃). Transfer 25 mL of synthetic hard water stock by pipette to a 1000 mL volumetric flask and dilute to volume with water at 20°C. Mix the solution thoroughly (see Note 3).

NOTE 3—It is recommended that the total hardness be checked in accordance with Test Method MT-73, CIPAC 1, EDTA titration.⁵ An alternate method is provided in Test Method D 1126, where the value is represented as CaCO₃. A value within ± 5 % of the nominal hardness value is acceptable.

² *Annual Book of ASTM Standards*, Vol 11.01.

³ *Annual Book of ASTM Standards*, Vol 15.05.

⁴ *Annual Book of ASTM Standards*, Vol 11.05.

Reagent Chemicals, American Chemical Society Specifications, American Chemical Society, Washington, DC. For suggestions on the testing

⁵ "Analysis of reagents not listed by the American Chemical Society, see *Analar Standards for Laboratory Chemicals*, BDH Ltd., Poole, Dorset, U.K.; Technical and the Formulated Pesticides," *United States Pharmacopeia and National Formulary: CIPAC Handbook*, U.S. Pharmacopeial Convention, Inc. (USPC), Rockville, MD, Vol 1, compiled by Ashworth, R de B., Henriet, J., Lovett, J.F., Collaborative International Pesticide Analytical Council Ltd., Great Britain, 1970.

6.5 *pH Reference Solution*, pH 4.00.

6.6 *pH Reference Solution*, pH 7.00.

6.7 *pH Reference Solution*, pH 10.00.

6.8 *0.05 N Sodium Hydroxide*, prepared according to Practice E 200, or the equivalent.

6.9 *0.05 N Sulfuric Acid*, prepared according to Practice E 200, or the equivalent.

6.10 *0.05 N Hydrochloric Acid*, prepared according to Practice E 200, or the equivalent.

TEST METHOD A—MEASUREMENT OF pH CHANGE OF STANDARD HARD WATER

7. Procedure

~~7.1 Turn on the~~

~~7.1~~ The pH meter should be turned on and allowed to warm up according to manufacturers instructions.

~~7.2~~ Calibrate the pH meter according to manufacturers directions using the pH 4.00 and pH 7.00 reference standards for an acidifier or the pH 7.00 and pH 10.00 reference standards for a basic blend (see Note 4). Place the standards in a 250 mL beaker and stir with the magnetic stirrer and stir bar, while the measurements are being made. Care should be taken to thoroughly wash the electrodes with water between measurements.

~~7.3~~ Place a clean stir bar and 100 mL of standard hard water (see Note 4) in a 250 mL beaker. Begin stirring with the magnetic stirrer. After carefully rinsing the electrodes, place them in the water and measure and record the pH, which should be between 5.5 and 7.

NOTE 4—Other or additional reference standards may be used (see Test Method E 70) as long as the pHs measured are not outside the range of the standards used.

~~7.3~~ Place a clean stir bar and 100 mL of standard hard water (see Note 5) in a 250 mL beaker. Begin stirring with the magnetic stirrer. After carefully rinsing the electrodes, place them in the water and measure and record the pH, which should be between 5.5 and 7.

NOTE 5—Other water samples can be used if appropriate. In this case the initial pH may not be between 5.5 and 7. For specific applications the actual water to be used is preferred.

~~7.4~~ Add 2 mL of the adjuvant to be tested to the beaker and allow to stir for ~~5 minutes.~~ min. Other dilutions may be used if specified on the product label.

7.5 Record the new pH value.

7.6 Remove the electrodes from the solution, clean them, and store them properly. Dispose of the solution properly.

8. Report

8.1 Report the initial and final pH values (see Note ~~5).~~ 6).

NOTE ~~5~~ 6—For an acidifier the final pH should be lower than the initial. For a basic blend it should be higher.

9. Precision and Bias

9.1 The precision and bias of this test method has not been determined.

TEST METHOD B—~~ACIDIFICATION B—~~BUFFER CAPACITY

10. Procedure

10.1 Turn on and warm up the pH meter according to manufacturers directions.

10.2 Calibrate the pH meter according to manufacturers directions using the pH 4.00 and pH 7.00 reference standards for an acidifier or the pH 4.00 and pH 10.00 reference standards for a basic blend (see Note 7). The standards should be placed in a 250 mL beaker and stirred with the magnetic stirrer and stir bar, while the measurements are being made. Care should be taken to thoroughly wash the electrodes with water between measurements.

~~10.3~~ Place a clean stir bar and 100 mL of standard hard water (see Note 6) in a 250 mL beaker. Begin stirring with the magnetic stirrer. After carefully rinsing the electrodes, place them in the water and measure and record the pH, which should be between 5.5 and 7.

NOTE ~~6~~ 7—Other or additional reference standards may be used (see Test Method E 70) as long as the pH's measured are not outside the range of the standards used.

~~10.3~~ Place a clean stir bar and 100 mL of standard hard water (see Note 8) in a 250 mL beaker. Begin stirring with the magnetic stirrer. After carefully rinsing the electrodes, place them in the water and measure and record the pH, which should be between 5.5 and 7.

NOTE 8—Other water samples can be used if appropriate. In this case the initial pH may not be between 5.5 and 7. For specific applications the actual water to be used is preferred.

10.4 For an acidifier fill the burette with 0.05 *N* sodium hydroxide. For a basic blend fill the burette with 0.05 sodium hydroxide and record *N* hydrochloric acid or 0.05 *N* sulfuric acid. Record the initial volume.

10.5 Weigh 2 ± 0.1 g to the nearest 0.01 grams of the adjutant to be tested (see Note 7) 9) into the beaker and allow to stir for 5 min.

NOTE 7—If necessary for a particular product, the sample size should be adjusted to give a titration volume between 5 and 40 mLs.

10.6 Record the pH versus volume titration curve. The curve should be measured carefully in the region of the initial water pH.

10.7 From the titration curve determine the volume of 0.05 *N* sodium hydroxide solution necessary titrant required to return the sample pH to its initial value.

11. Calculation

11.1 Calculate the number of milliequivalents of acid (for an acidifier) or base (for a basic blend) per gram of product. Multiply the titration volume times the concentration of the NaOH titrant solution. Divide this result by the product weight:

$$\text{milliequivalents of acid-gram of adjutant} = \frac{(\text{titration volume (mL)} * 0.05 \text{ meq/mL})}{\text{sample weight (grams)}} \quad (1)$$

$$\frac{\text{Milliequivalents of acid or base}}{\text{Gram of Adjuvant}} = \frac{(\text{Titration Volume (mL)} * 0.05 \text{ (meq/mL)})}{\text{Sample Weight (g)}} \quad (1)$$

12. Report

12.1 Report the number of milliequivalents of acid (for an acidifier) or base (for a basic blend) per gram of adjutant.

13. Precision and Bias

13.1 The precision and bias of this test method has not been determined.

14. Keywords

14.1 acidifier; adjutant; agricultural adjutant; basic blend; pH

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